



Selvam College of Technology



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Name of the Bundle	Proficient Bundle V1	Subject	Competitive Exams Training (Science & Geography)
Topic	Periodic Classification	Last updated on	23 August 2024

1. The most electronegative element in the periodic table is ____

- a) Fluorine
- b) Chlorine
- c) Nitrogen
- d) None of the above

Ans: a) Fluorine

2. Which metalloid in the carbon group is chemically similar to its group neighbours, tin and silicon?

- a) Germanium
- b) Flerovium
- c) Lead
- d) Arsenic

Ans: a) Germanium

3. Name the group in the Modern Periodic Table which has elements that are all gases.

- a) Group 13
- b) Group 10
- c) Group 18
- d) Group 12

Ans: c) Group 18

4. In terms of ionisation enthalpy, H resembles which of the group's elements more?

- a) Group 1
- b) Group 17
- c) Group 18
- d) Group 16

Ans: b) Group 17

5. In the Modern Periodic Table, atomic size increases down the group because _____

- a) Nuclear charge increases
- b) Tendency to lose electrons decreases
- c) Number of shells increases
- d) None of the above

Ans: c) Number of shells increases

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6. In which group of the modern periodic table are halogens placed?

- a) 16th
- b) 18th
- c) 17th
- d) 19th

Ans: c) 17th

7. What number of elements are provided in the modern periodic table?

- a) 108
- b) 100
- c) 94
- d) 118

Ans: d) 118

8. Which of the following is incorrect about the modern periodic table?

- a) Modern periodic table has 15 columns.
- b) From left to right a period, each element has one more proton than the element before it.
- c) Columns of the modern table are called groups.
- d) Rows of the modern periodic table are called periods.

Ans: a) Modern periodic table has 15 columns.

9. Elements of the 2nd period are known as _____.

- a) Typical element
- b) Bridge element
- c) Transitional element
- d) Normal element

Ans: b) Bridge element

10. Which of the following elements has zero valency?

- a) Lithium
- b) Beryllium
- c) Helium
- d) Fluorine

Ans: c) Helium



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11. The most electropositive halogen is

- a) Fluorine
- b) Chlorine
- c) Bromine
- d) Iodine

Ans: d) Iodine

12. Which of the following elements are included in the same group in the periodic table?

- a) Ne, Ar, Na, Mg
- b) He, Li, Mg, K
- c) H, Li, Mg, Ca
- d) H, Li, Na, K

Ans: d) H, Li, Na, K

13. Modern periodic law states that the physical and chemical properties of elements are the periodic functions of their _____.

- a) atomic numbers
- b) atomic masses
- c) similarities
- d) anomalies

Ans: a) atomic numbers

14. What is the position of the metalloids in the modern periodic table?

- a) Left of the table
- b) Right of the table
- c) Middle of the table
- d) Bottom of the table

Ans: b) Right of the table

15. The valency of noble gas elements is _____

- a) 3
- b) 1
- c) 0
- d) 2

Ans: c) 0



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16. Who among the following made the first observation on Platinum as a catalyst and discovered similar triads of elements that led to the development of the Periodic Table of elements?

- a) Dmitri Mendeleev
- b) Hans Christian Oersted
- c) Dobereiner
- d) Michael Faraday

Ans: c) Dobereiner

17. Among the given elements, which comes first in the periodic table of elements?

- a) Sodium
- b) Neon
- c) Potassium
- d) Silicon

Ans: b) Neon

18. Noble gases belong to which of the following groups of the periodic table?

- a) 16
- b) 18
- c) 13
- d) 15

Ans: b) 18

19. How many elements are there in the periodic table such that for every element name there is a single alphabet that symbolically represents it ?

- a) 10
- b) 17
- c) 7
- d) 14

Ans: d) 14

20. Which of the following does NOT increase while moving down the group of the periodic table?

- a) Number of shells in the element
- b) Metallic character
- c) Valency
- d) Atomic radius

Ans: c) Valency



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21. Which element is the first element in the periodic table?

- a) Neon
- b) Hydrogen
- c) Oxygen
- d) Helium

Ans: b) Hydrogen

22. According to Mendeleev's periodic law, the elements were arranged in the periodic table in the order of

- a) Increasing atomic masses
- b) Decreasing atomic number
- c) Increasing atomic number
- d) Decreasing atomic masses

Ans: a) Increasing atomic masses

23. Na, Mg, Al, Si, P, S, Cl and Ar belong to the ____ period of the modern periodic table.

- a) Third
- b) Second
- c) Fourth
- d) First

Ans: a) Third

24. The modern periodic table has vertical columns known as _____

- a) Columns
- b) Periods
- c) Table
- d) Groups

Ans: d) Groups

25. Non-metals like sulphur and chlorine are found on the _____ of the periodic table.

- a) Middle
- b) Right hand side
- c) Zigzag
- d) Left hand side

Ans: b) Right hand side



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26. While moving down in a group, the metallic character _____

- a) Increases
- b) Decreases
- c) Remains constant
- d) First increases then decreases

Ans: a) Increases

27. What is the maximum number of electrons that can be accommodated in a k shell?

- a) 2 electrons
- b) 8 electrons
- c) 18 electrons
- d) 32 electrons

Ans: a) 2 electrons

28. In which of the following liquids can sodium metal be stored safely?

- a) Water
- b) Dilute sulphuric acid
- c) Kerosene
- d) Nitric acid

Ans: c) Kerosene

29. The more reactive member in halogen is _____

- a) Chlorine
- b) Fluorine
- c) Iodine
- d) Sulphur

Ans: b) Fluorine

30. The elements in the first group of s- block are also known as _____

- a) Alkali metals
- b) Alkaline earth metals
- c) Halogens
- d) Noble gases

Ans: a) Alkali metals



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31. which among the following has lowest electron affinity?

- a) Fluorine
- b) Chlorine
- c) Bromine
- d) Argon

Ans: d) Argon

32. Element having $Z = 68$ is related to _____

- a) s- block
- b) p- block
- c) d- block
- d) f- block

Ans: d) f- block

33. Which year is declared as the International Year of the Periodic Table of Chemical Elements to commemorate the 150th birth anniversary of the periodic table of chemical elements ?

- a) 2014
- b) 2015
- c) 2019
- d) 2016

Ans: c) 2019

34. How many rare earth elements are present in the periodic table?

- a) 17
- b) 19
- c) 21
- d) 24

Ans: a) 17

35. What was the position of noble gases in Mendeleev's periodic table?

- a) Noble gases, being very inert, were not necessarily put in the periodic table.
- b) Noble gases were there in the seventh group.
- c) Mendeleev divided all elements into 8 groups, so there was no place for noble gases.
- d) Noble gases were not discovered by that time.

Ans: d) Noble gases were not discovered by that time.



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36. Electronegativity _____ from left to right in a row in the periodic table and electropositivity _____ from left to right in a row in the periodic table.

- a) Decreases, increases
- b) Decreases, decreases
- c) Increases, increases
- d) Increases, decreases

Ans: d) Increases, decreases

37. Elements in the Modern Periodic Table are arranged in _____ periods.

- a) 6
- b) 8
- c) 5
- d) 7

Ans: d) 7

38. An element of atomic number 16 will belong to the _____ period of the periodic table.

- a) 5th
- b) 6th
- c) 3rd
- d) 4th

Ans: c) 3rd

39. How many groups and periods are present in the Modern Periodic Table?

- a) 8 groups and 6 periods
- b) 18 groups and 9 periods
- c) 18 groups and 7 periods
- d) 7 groups and 10 periods

Ans: c) 18 groups and 7 periods

40. Which of the following inert gases is placed in period 4 of the periodic table?

- a) Kr
- b) Rn
- c) Xe
- d) Ar

Ans: a) Kr



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41. The only block which consists of metals, nonmetals and metalloids in a periodic table is

- a) S- Block
- b) P- Block
- c) D- Block
- d) F- Block

Ans: b) P- Block

42. Who predicted the existence of some yet to be discovered elements on the basis of gaps in his periodic table ?

- a) Mendeleev
- b) Dobereiner
- c) Newlands
- d) Mosely

Ans: a) Mendeleev

43. Who has developed the Modern Periodic table?

- a) Julius Meyer
- b) John Newlands
- c) Dmitri Mendeleev
- d) Henry Moseley

Ans: d) Henry Moseley

44. Pyrogens are arranged in which block of the periodic table?

- a) f- block
- b) d- block
- c) p- block
- d) s- block

Ans: c) p- block

45. _____ is the only non-metal present in Group 1 of the Modern Periodic Table.

- a) Cobalt
- b) Rubidium
- c) Potassium
- d) Hydrogen

Ans: d) Hydrogen



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46. Which of the following elements was the last element in Newland's Law of Octaves?

- a) Rubidium
- b) Thorium
- c) Bromine
- d) Hydrogen

Ans: b) Thorium

47. All acids generate _____ gas reacting with metals.

- a) Hydrogen
- b) Chlorine
- c) Nitrogen
- d) Oxygen

Ans: a) Hydrogen

48. Each water molecule contains _____ hydrogen atoms and one oxygen atom.

- a) Two
- b) Three
- c) One
- d) Four

Ans: a) Two

49. Except Helium, all noble gases, how many electrons in the outermost shell?

- a) 10
- b) 6
- c) 4
- d) 8

Ans: d) 8

50. The tenth element in the Newland's periodic classification resembles the ____

- a) First
- b) Fourth
- c) Ninth
- d) Third

Ans: d) Third



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51. Which of the following scientists has done some calculations for the fifth state of matter?

- a) CV Raman
- b) Satyendra Nath Bose
- c) Vikram Sarabhai
- d) Homi Bhabha

Ans: b) Satyendra Nath Bose

52. Which is the non metal laced at the left side of the Modern Periodic Table?

- a) Hydrogen
- b) Carbon
- c) Neon
- d) Helium

Ans: a) Hydrogen

53. In the Modern Periodic Table, there is an anomaly when it comes to the position of..... because it can be placed either in group 1 or group 17 in the first period.

- a) He
- b) H
- c) Li
- d) Be

Ans: b) H

54. Non-metals are generally more electronegative due to their

- a) Large number of electrons
- b) Smaller atomic radii
- c) Smaller atomic number
- d) Smaller ionization energy

Ans: b) Smaller atomic radii

55. Oxides of non-metals are usually _____

- a) Acidic
- b) Neutral
- c) Less reactive
- d) Basic

Ans: a) Acidic



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56. Which element in the periodic table forms the maximum number of compounds?

- a) Hydrogen
- b) Oxygen
- c) Sulphur
- d) Carbon

Ans: d) Carbon

57. Name the noble gas which is placed in the third period and eighth group of the Modern Periodic Table.

- a) Helium
- b) Argon
- c) Neon
- d) Krypton

Ans: b) Argon

58. _____ elements are present in the sixth period of the modern periodic table.

- a) 32
- b) 18
- c) 33
- d) 8

Ans: a) 32

59. According to _____, the properties of elements are a periodic function of their atomic masses.

- a) Newland Law of Octaves
- b) Mendeleev's periodic law
- c) Dobereiner's Law of Triads
- d) Newton's Law

Ans: b) Mendeleev's periodic law

60. The element of the Lanthanide series having the atomic number 58 is _____

- a) Cerium
- b) Lanthanum
- c) Thorium
- d) Strontium

Ans: a) Cerium



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61. _____ has the greatest electron affinity.

- a) Br
- b) Cl
- c) I
- d) F

Ans: b) Cl

62. The maximum number of electrons that can be accommodated in M shell is:

- a) 8
- b) 12
- c) 18
- d) 32

Ans: c) 18

63. Which of the following elements replaced eka-Aluminium in Mendeleev's Periodic Table?

- a) Scandium
- b) Gallium
- c) Titanium
- d) Germanium

Ans: b) Gallium

64. Metals are placed on which side of the Modern Periodic Table?

- a) Top row
- b) Left side
- c) Right side
- d) Bottom row

Ans: b) Left side

65. In the Modern Periodic Table elements present in the same period will have the same

- a) Atomic weight
- b) Number of shells
- c) Atomic number
- d) Valence electrons

Ans: b) Number of shells